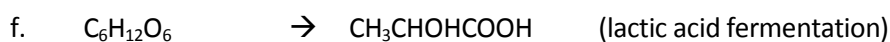
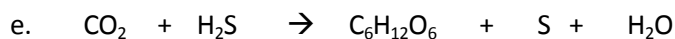
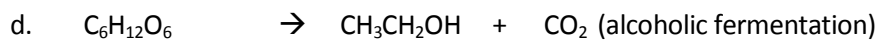
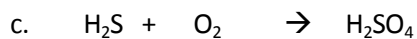
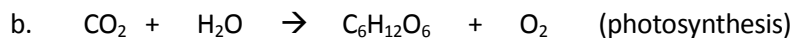
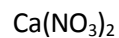
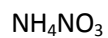


Stoichiometry Practice

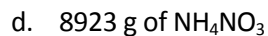
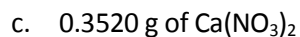
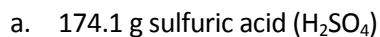
1. Each of the following equations represents a metabolic activity carried out by various species of living organisms. Balance each equation.



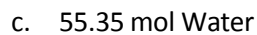
2. Calculate the molar mass for each of the following:



3. How many moles are in each of the following samples?

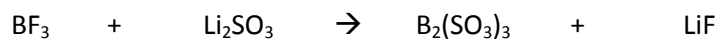


4. How many grams are in each of the following samples?



Stoichiometry Practice

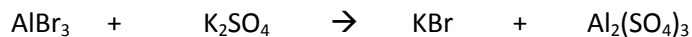
5. Answer the following questions given this equation:



- Balance the equation
 - Calculate molar masses for each molecule
- c. If 37.29g BF_3 is reacted with excess Li_2SO_3 , how many grams of $\text{B}_2(\text{SO}_3)_3$ could be formed by the reaction? How many grams of LiF ?
- d. If 189.4g of LiF are desired, how many grams of Li_2SO_3 are required? How many grams BF_3 ?
- e. If 71.90g BF_3 are reacted with 77.01g Li_2SO_3 :
- Which reactant will be the limiting reagent?
 - How many grams of LiF can be made, given this limit?
 - How many grams of the excess reagent will remain leftover?
 - If, in the laboratory, the actual mass of LiF produced was 56.22g, what is the percent yield for the reaction?

Stoichiometry Practice

6. Answer the following questions given this equation:



- a. Balance the equation
- b. Calculate molar masses for each molecule

- c. If 500.g of AlBr_3 are reacted with excess K_2SO_4 , how many grams of KBr are possible? How many grams $\text{Al}_2(\text{SO}_4)_3$ possible?

- d. If 190.0g KBr is desired, how many grams K_2SO_4 would be needed? How many grams AlBr_3 ?

- e. If 91.97g AlBr_3 are reacted with 152.01g K_2SO_4 :
 - i. Which reactant will be the limiting reagent?

 - ii. How many grams of KBr can be made, given this limit?

 - iii. How many grams of the excess reagent will remain leftover?

 - iv. If, in the laboratory, the percent yield of this reaction is 65.2%, how many grams of KBr will actually be produced after completing the reaction?