

Balancing Practice

1. Balance the Following:

- a. $\underline{\quad} \text{Na}_3\text{PO}_4 + \underline{\quad} \text{KOH} \rightarrow \underline{\quad} \text{NaOH} + \underline{\quad} \text{K}_3\text{PO}_4$
- b. $\underline{\quad} \text{MgF}_2 + \underline{\quad} \text{Li}_2\text{CO}_3 \rightarrow \underline{\quad} \text{MgCO}_3 + \underline{\quad} \text{LiF}$
- c. $\underline{\quad} \text{P}_4 + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{P}_2\text{O}_3$
- d. $\underline{\quad} \text{RbNO}_3 + \underline{\quad} \text{BeF}_2 \rightarrow \underline{\quad} \text{Be}(\text{NO}_3)_2 + \underline{\quad} \text{RbF}$
- e. $\underline{\quad} \text{AgNO}_3 + \underline{\quad} \text{Cu} \rightarrow \underline{\quad} \text{Cu}(\text{NO}_3)_2 + \underline{\quad} \text{Ag}$
- f. $\underline{\quad} \text{CF}_4 + \underline{\quad} \text{Br}_2 \rightarrow \underline{\quad} \text{CBr}_4 + \underline{\quad} \text{F}_2$
- g. $\underline{\quad} \text{HCN} + \underline{\quad} \text{CuSO}_4 \rightarrow \underline{\quad} \text{H}_2\text{SO}_4 + \underline{\quad} \text{Cu}(\text{CN})_2$
- h. $\underline{\quad} \text{GaF}_3 + \underline{\quad} \text{Cs} \rightarrow \underline{\quad} \text{CsF} + \underline{\quad} \text{Ga}$
- i. $\underline{\quad} \text{BaS} + \underline{\quad} \text{PtF}_2 \rightarrow \underline{\quad} \text{BaF}_2 + \underline{\quad} \text{PtS}$
- j. $\underline{\quad} \text{N}_2 + \underline{\quad} \text{H}_2 \rightarrow \underline{\quad} \text{NH}_3$
- k. $\underline{\quad} \text{NaF} + \underline{\quad} \text{Br}_2 \rightarrow \underline{\quad} \text{NaBr} + \underline{\quad} \text{F}_2$
- l. $\underline{\quad} \text{Pb}(\text{OH})_2 + \underline{\quad} \text{HCl} \rightarrow \underline{\quad} \text{H}_2\text{O} + \underline{\quad} \text{PbCl}_2$
- m. $\underline{\quad} \text{AlBr}_3 + \underline{\quad} \text{K}_2\text{SO}_4 \rightarrow \underline{\quad} \text{KBr} + \underline{\quad} \text{Al}_2(\text{SO}_4)_3$
- n. $\underline{\quad} \text{CH}_4 + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{CO}_2 + \underline{\quad} \text{H}_2\text{O}$
- o. $\underline{\quad} \text{Na}_3\text{PO}_4 + \underline{\quad} \text{CaCl}_2 \rightarrow \underline{\quad} \text{NaCl} + \underline{\quad} \text{Ca}_3(\text{PO}_4)_2$
- p. $\underline{\quad} \text{K} + \underline{\quad} \text{Cl}_2 \rightarrow \underline{\quad} \text{KCl}$
- q. $\underline{\quad} \text{Al} + \underline{\quad} \text{HCl} \rightarrow \underline{\quad} \text{H}_2 + \underline{\quad} \text{AlCl}_3$
- r. $\underline{\quad} \text{S} + \underline{\quad} \text{HNO}_3 \rightarrow \underline{\quad} \text{H}_2\text{SO}_4 + \underline{\quad} \text{NO}_2 + \underline{\quad} \text{H}_2\text{O}$
- s. $\underline{\quad} \text{Cu} + \underline{\quad} \text{HNO}_3 \rightarrow \underline{\quad} \text{Cu}(\text{NO}_3)_2 + \underline{\quad} \text{NO} + \underline{\quad} \text{H}_2\text{O}$